UNITED NATIONS SCHOOL I.E.D.

PEI: COMPREHENSIVE TRAINING OF COMPETENT ENTREPRENEURIAL LEADERS, WITH DEMOCRATIC, TECHNOLOGICAL, CULTURAL AND SPORTS PRINCIPLES

MOTTO: "EDUCATION, SCIENCE, CULTURE AND SPORT TO TRANSCEND"

PREPARATION WORKSHOP FOR THE THIRD PERIOD

CHEMISTRY

SEVENTH GRADE

TEACHER HEISEL QUESADA

The preparation workshop must be carried out in the Chemistry notebook as a requirement to take the competency test

Delivery date: november 8

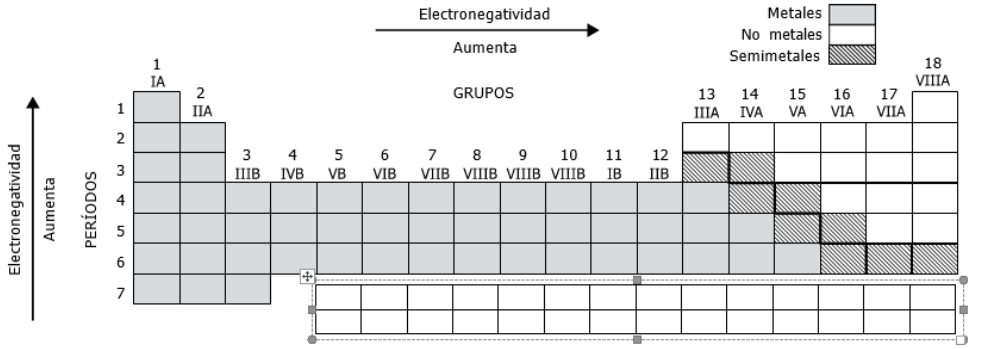
ANSWER QUESTIONS 1 TO 3 WITH THE FOLLOWING INFORMATION

A student read that the researcher Joseph Priestley, in 1771, carried out the following experiment: he put a mouse inside a transparent glass box that prevented air from entering from the outside and after a short time the mouse died. He then placed a lit candle in the same clear glass box and after a short time the candle went out.

1. Write a conclusion for Priestley's experiment
2. What role does oxygen play in the experiment in each case?
3. Check out Joseph Priestley's biography and what contributions he had to Chemistry

ANSWER QUESTIONS 4 TO 10 WITH THE FOLLOWING INFORMATION

In the periodic table, elements are arranged into groups according to similar physical and chemical properties. Elements are classified as metals, nonmetals, and semimetals. The following figure shows the location of metals, nonmetals, and semimetals on the periodic table.



Las siguientes fichas muestran información sobre las propiedades físicas y químicas de cuatro elementos del cuarto período.

X	*Q*	*R*	*T*
• Electronegatividad = 0,8 • Es maleable. • Presenta alta conductividad. • Electrones de valencia = 1	• Electronegatividad = 2,8 • No es ductil. • Presenta baja conductividad. • Electrones de valencia = 7	• Electronegatividad = 1,5 • Tiene brillo. • Presenta alta conductividad. • Electrones de valencia = 5	• Electronegatividad = 1,9 • Sólido maleable. • Presenta alta conductividad. • Electrones de valencia = 6

4. Write down the relationship between electronegativity and the groups and periods of the periodic table
5. From the diagram make a list of the metals, non-metals and semimetals of the periodic table
6. Organize the following elements from lowest to highest electronegativity: bromine, chlorine, sodium, radium, beryllium, sulfur, technetium, iodine, calcium, iron, cobalt, rhodium, nickel, thorium, astate, aluminum, gallium, tin, francium, and silver
7. What properties do the elements X, Q, R, T have?
8. Organize the elements X, Q, R, T according to their electronegativity and write whether they are metals, nonmetals or semimetals
9. Consult the meaning of the terms: ductile, conductivity, malleable, valence electrons and write their relationship with the chemical elements
10. Write down the importance of electronegativity for chemical elements